The equipartition theorem says that the mean energy in thermal equilibrium is $(q/2)kT$, where $q$ is the number of 'quadratic degrees of freedom'. A quadratic degree of freedom is one on which the energy depends quadratically. For instance there may be energy terms $\frac{1}{2}mv_x^2$, $\frac{1}{2}kx^2$, $\frac{1}{2}ky^2$, $\frac{1}{2}I\omega_x^2$ associated with degrees of freedom $v_x$, $x$, $y$, and $\omega_x$. (You can prove the equipartition theorem if you assume a Boltzmann distribution of energies and finding the mean energy).

For a gas of free noninteracting atoms, there are three quadratic degrees of freedom for each atom, $v_x$, $v_y$ and $v_z$, so $q = 3$, leading to $pV = (3/2)kT$. (Although there are three more degrees of freedom associated with position, they don't count, because the energy doesn't depend on position). What is $q$ for each of the following have? In each case, try to describe, or sketch, the degrees of freedom.

$$E = \tfrac{1}{2}mv_x^2 + \ldots + \tfrac{1}{2}kx^2 + \ldots + \tfrac{1}{2}I_1\omega_1^2 + \ldots \tfrac{1}{2}k_r r^2 + \ldots + \tfrac{1}{2}m_r \dot{v}^2 + \ldots$$

(a) An atom in a 3D parabolic potential well (eg, in an 'atom trap' in the Phys-Astr basement)

6 – why?　　$V_x$　$V_y$　$V_z$　$x$　$y$　$z$　$\left[ E = \tfrac{1}{2}m|\vec{v}|^2 + \tfrac{1}{2}k|\vec{r}|^2 \right]$

(b) An atom in a solid

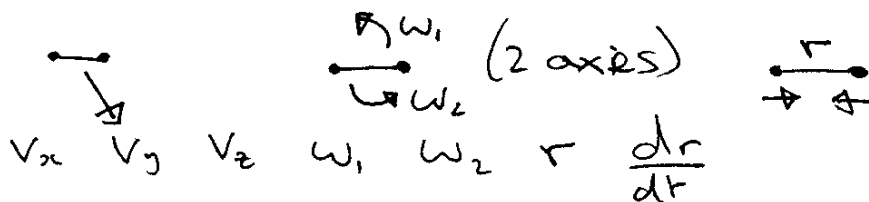(6) – as for an atom in a trap. It is confined to a position by its neighbors.

(c) Two free O atoms

(6)　$V_{x1}$　$V_{y1}$　$V_{z1}$　　$V_{x2}$　$V_{y2}$　$V_{z2}$

(d) A free diatomic molecule (eg $O_2$) – including vibrations and rotations!　　(7)

$\omega_1$　(2 axes)

$V_x$　$V_y$　$V_z$　$\omega_1$　$\omega_2$　$r$　$\dfrac{dr}{dt}$
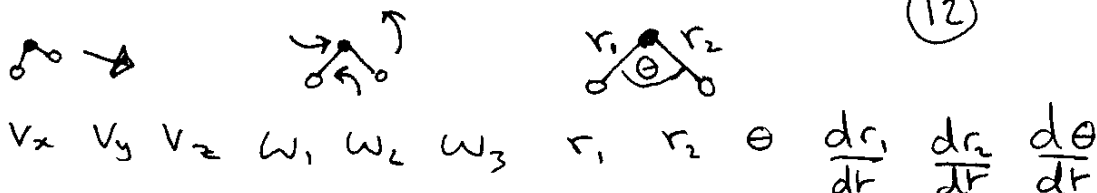
(e) Two free H atoms and one free O atom.

(9) = 3 + 3 + 3 !

(f) A free water molecule ($H_2O$)

(12)

$V_x$　$V_y$　$V_z$　$\omega_1$　$\omega_2$　$\omega_3$　$r_1$　$r_2$　$\theta$　$\dfrac{dr_1}{dt}$　$\dfrac{dr_2}{dt}$　$\dfrac{d\theta}{dt}$

(f) A water molecule in liquid water (if you can do that, you will win a Nobel prize some day)

Lots.